

M.H. Khan 9708421577, 21-4-2017
 Topic - Equivalent wt and Prob. Molar coll

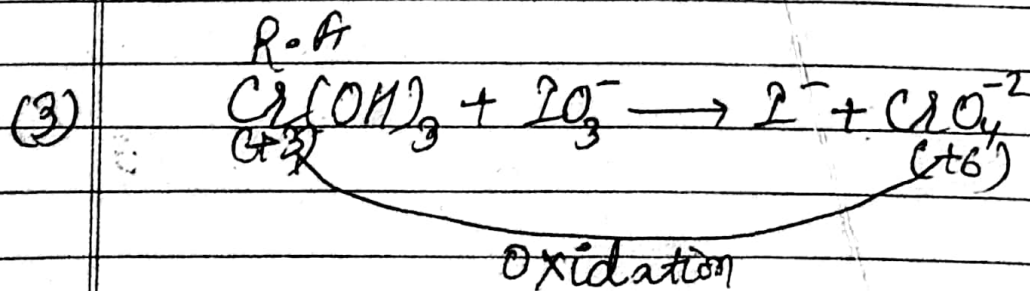
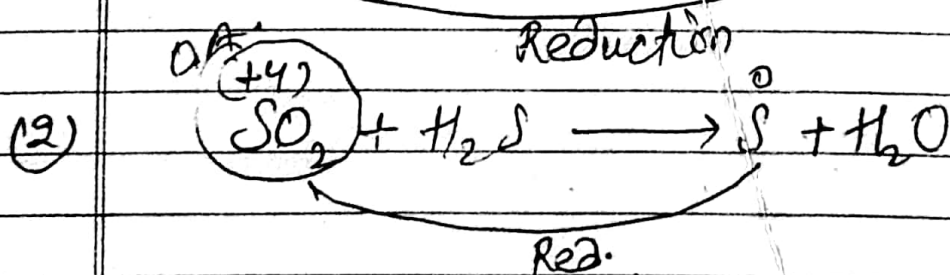
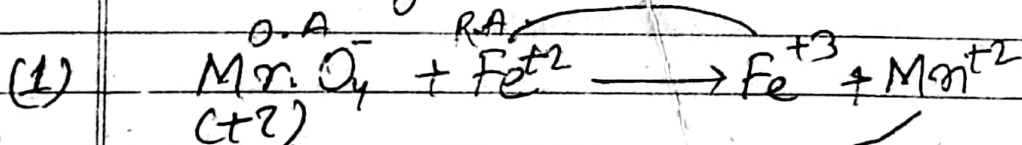
(i) Li is most powerful reducing agent
(ii) NaH , H_2S , CaH_2
(-1), (-2), (-1)

xxx SO_2 , H_2O_2 and HNO_2 behave as a oxidant and reductant.

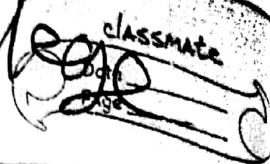
17-03

Ex. 1 and 2 All.

Q. Calculate the oxidant and Reductant in following rxn.



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eq wt + Prob. mawari college
XII



Application of Oxidation no.

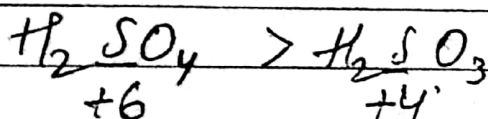
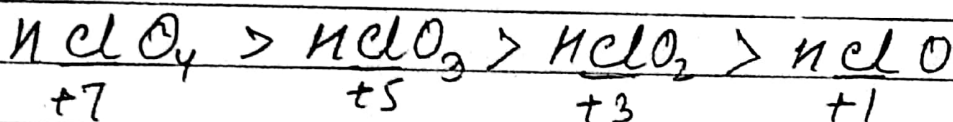
1. Determination of Molecular formula of comp by oxidation no.

Ex: A B C.
Ox. no. +2 +1 -2.

- (i) $(AB)_2C$ $2(+1) + 2(+1) - 2 \neq 0$
 (ii) $A_2B_2C_3$ $2(+2) + 2(+1) + 3(-2) = 0$
 (iii) $(AB)_3C$
 (iv) None.

2. Acidic nature is proportional oxidation no. of central atom, all same.

Acidic nature $\propto$ Ox. no.



3. Equivalent wt.

$$\text{Eq. wt} = \frac{\text{At. mass / Molar mass}}{n\text{-factor}}$$

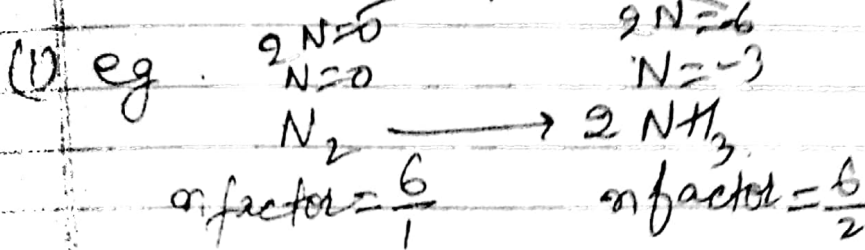
- A) Determination of equivalent wt. of intermolecular redox rxn (Oxidation or Reduction)

$$n\text{-factor} = \frac{\text{Loss of } e^- / \text{gain of } e^-}{\text{Stoichiometric coefficient}}$$

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equi + Prob, marwari collage

XTH

n-factor = 6



$E = \frac{28}{6}$

$E = \frac{17}{3}$

